

Name: _____

Date: _____ Block: _____

Science 10 - Balancing Equations Review

Balance the following equations:

- 1) $\text{Mg} + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
- 2) $\text{Ag}_2\text{SO}_4 + \text{AlCl}_3 \rightarrow \text{AgCl} + \text{Al}_2(\text{SO}_4)_3$
- 3) $\text{BaCl}_2 + \text{MgSO}_4 \rightarrow \text{BaSO}_4 + \text{MgCl}_2$
- 4) $\text{C}_{12}\text{H}_{22}\text{O}_{11} \rightarrow \text{C} + \text{H}_2\text{O}$
- 5) $\text{KI} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{PbI}_2 + \text{KNO}_3$
- 6) $\text{H}_2\text{SO}_4 + \text{NH}_4 \rightarrow (\text{NH}_4)_2\text{SO}_4 + \text{H}_2$
- 7) $\text{K}_{(s)} + \text{Cl}_{2(g)} \rightarrow \text{KCl}_{(s)}$
- 8) $\text{Ca}_{(s)} + \text{H}_2\text{O}_{(l)} \rightarrow \text{Ca}(\text{OH})_{2(s)} + \text{H}_{2(g)}$
- 9) $\text{CuSO}_{4(s)} + \text{Fe}_{(s)} \rightarrow \text{Fe}_2(\text{SO}_4)_{3(s)} + \text{Cu}_{(s)}$
- 10) $\text{Br}_{2(g)} + \text{KI}_{(s)} \rightarrow \text{KBr}_{(s)} + \text{I}_{2(g)}$

Write the chemical equation for the following word equations, then balance.

- 1) lead II nitrate + potassium iodide \rightarrow lead II iodide + potassium nitrate

- 2) magnesium nitrate + sodium carbonate \rightarrow magnesium carbonate + sodium nitrate

- 3) chlorine gas reacts with sodium bromide to give sodium chloride and bromine gas.

- 4) In solution, silver nitrate reacts with sodium chloride to give silver chloride and sodium nitrate.

- 5) Copper II oxide combined with hydrogen gas produces copper metal and water.