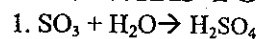


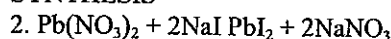
WRITE THE FORMULA FOR EACH MATERIAL CORRECTLY AND THEN BALANCE THE EQUATION. THERE ARE SOME REACTIONS THAT REQUIRE COMPLETION. FOR EACH REACTION TELL WHAT TYPE OF REACTION IT IS.

1. sulfur trioxide and water combine to make sulfuric acid.
2. lead II nitrate and sodium iodide react to make lead iodide and sodium nitrate.
3. calcium fluoride and sulfuric acid make calcium sulfate and hydrogen fluoride (Hydrofluoric acid)
4. calcium carbonate will come apart when you heat it to leave calcium oxide and carbon dioxide.
5. ammonia gas when it is pressed into water will make ammonium hydroxide.
6. sodium hydroxide neutralizes carbonic acid
7. zinc sulfide and oxygen become zinc oxide and sulfur.
8. lithium oxide and water make lithium hydroxide
9. aluminum hydroxide and sulfuric acid neutralize to make water and aluminum sulfate.
10. sulfur burns in oxygen to make sulfur dioxide.
11. barium hydroxide and sulfuric acid make water and barium sulfate.
12. aluminum sulfate and calcium hydroxide become aluminum hydroxide and calcium sulfate.
13. copper metal and silver nitrate react to form silver metal and copper II nitrate.
14. sodium metal and chlorine react to make sodium chloride.
15. calcium phosphate and sulfuric acid make calcium sulfate and phosphoric acid.
16. phosphoric acid plus sodium hydroxide.
17. propane burns (with oxygen)
18. zinc and copper II sulfate yield zinc sulfate and copper metal
19. sulfuric acid reacts with zinc
20. steam methane to get hydrogen and carbon dioxide
21. calcium oxide and aluminum make aluminum oxide and calcium
22. chlorine gas and sodium bromide yield sodium chloride and bromine

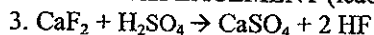
ANSWERS TO EQUATIONS



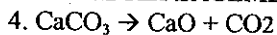
SYNTHESIS



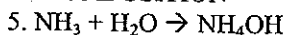
DOUBLE REPLACEMENT (lead II iodide precipitates)



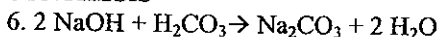
DOUBLE REPLACEMENT (calcium sulfate precipitates)



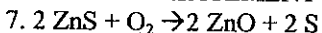
DECOMPOSITION



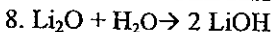
SYNTHESIS



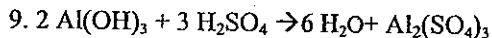
DOUBLE REPLACEMENT OR ACID-BASE NEUTRALIZATION



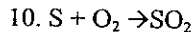
ANIONIC SINGLE REPLACEMENT



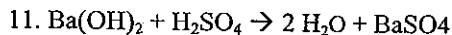
SYNTHESIS



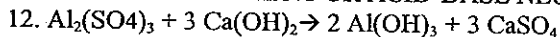
DOUBLE REPLACEMENT OR ACID-BASE NEUTRALIZATION



SYNTHESIS

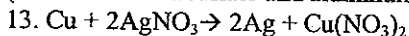


DOUBLE REPLACEMENT OR ACID-BASE NEUTRALIZATION

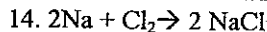


DOUBLE REPLACEMENT

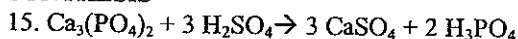
(BOTH calcium sulfate and aluminum hydroxide are precipitates.)



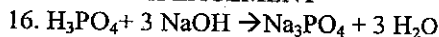
CATIONIC SINGLE REPLACEMENT



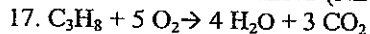
SYNTHESIS



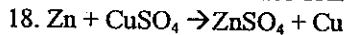
DOUBLE REPLACEMENT



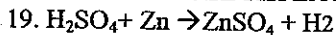
DOUBLE REPLACEMENT (NEUTRALIZATION)



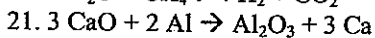
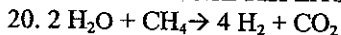
BURNING OF A HYDROCARBON



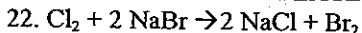
CATIONIC SINGLE REPLACEMENT



CATIONIC SINGLE REPLACEMENT



CATIONIC SINGLE REPLACEMENT



ANIONIC SINGLE REPLACEMENT